A Highly Specific Probe for Sensing Hydrogen Sulfide in Live Cells Based on Copper-Initiated Fluorogen with Aggregation-Induced Emission Characteristics

Xin Li1, Chengyu Yang2, Kai Wu2, Yongzhou Hu1, Yifeng Han2 and Steven H. Liang3

1. ZJU-ENS Joint Laboratory of Medicinal Chemistry, College of Pharmaceutical Sciences, Zhejiang University, Hangzhou 310058, China
2. The Key Laboratory of Advanced Textile Materials and Manufacturing Technology, Department of Chemistry, Zhejiang Sci-Tech University, Hangzhou, 310018, China.
3. Division of Nuclear Medicine and Molecular Imaging, Massachusetts General Hospital and Department of Radiology, Harvard Medical School, 55 Fruit St., Boston, MA 02114, USA.

Corresponding author s: Y. Han. Tel: 86 571 86843550, Email: hanyf@zstu.edu.cn; S.H. Liang Tel: 617 726 3404, Email: Liang.Steven@mgh.harvard.edu

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Abstract

Here we reported the first fluorescent probe with aggregation-induced emission characteristics, namely AIE-S, for the detection of hydrogen sulfide (H2S) in live cells. The detection system is selective for complicated biological application and the response is fast enough to complete within seconds. Moreover, the probe exhibits the unique advantage of being immune to aggregation-caused quenching which is a detrimental phenomenon limiting the application of most current available H2S fluorescent probes. The detection mechanism was investigated and postulated to be S2- initiated de-coordination and thereafter aggregation of the AIE-S complex.

Key words: hydrogen sulfide, fluorescent probe, cell imaging, aggregation-induced emission

Introduction

Hydrogen sulfide (H2S), notorious for the stink of rotten-eggs, was considered merely as a toxin with no beneficial physiological significance until the initial observations by Kimura’s group suggesting its biological relevance as a signaling molecule.1 Ever since then, many research efforts have been made to unveil its biological roles, which established soundly H2S as the third gasotransmitter after nitric oxide and carbon monoxide for its fundamental roles in human biology including its regulation and protection effects in the central nervous system, in the cardiovascular system, and in the inflammation process, etc.2,3 Hydrogen sulfide (pKₐ 6.755 at 37°C) is produced in mammals from cysteine in a series of reactions mainly catalyzed by cystathionine β-synthase (CBS) and cystathionine γ-lyase (CSE).4,5 Endogenous H2S is maintained at levels sufficient to modify cell function yet without causing poisonous effects in healthy conditions.4,5 However, abnormal H2S regulation has been associated with an array of diseases including hypertension,6 Alzheimer’s disease,7 Down’s syndrome8 and diabetes.9 H2S donors, on the other hand, have been reported to exhibit antihypertensive activity,10 antiproliferative effects,11 and neuromodulatory activity.12

Despite all the potential therapeutic implications,13 our understanding of the biological and pathological roles of this intriguing gas is still in its infancy. However, for any study of its biological functions, it is important to take the concentration at the site of action into first consideration. As such, it is crucial to have in place robust assays to monitor the
biosynthesis, transport and catabolism of H₂S in living cells, tissues, as well as whole organisms. Traditional methods for H₂S detection, including the methylene blue assay,14 sulfide-selective electrodes6 and gas chromatography15 are generally limited by their invasive and destructive nature to live organisms. On the other hand, small fluorescent probes confer the advantages of high sensitivity and good cell permeability, and have therefore attracted the most attention for sensing and visualizing analytes inside living cells. A number of probes have recently been developed for the fluorescence detection of H₂S in biological systems. These probes, judiciously designed by taking advantage of nucleophilic, reductive properties of H₂S or its high binding affinity towards copper, can react specifically with sulfides with a fluorescence output signal indicating the presence or even concentration of H₂S in biological systems.16-27 However, previous H₂S probes are based on traditional planar aromatic chromophores which are only fluorescent in dilute solutions since high concentrations usually induce aggregation-caused quenching (ACQ) effect resulting from the strong π π stacking interactions upon the aggregation formation.28 ACQ is generally considered harmful to practical applications because of the significant detection errors incurred. Unfortunately, traditional discotic fluorophores are usually highly hydrophobic and therefore easily aggregate in the aqueous biological contents. Although their hydrophilicity can be improved by the attachment of water-soluble functional groups, the resultant products are still vulnerable to ACQ in aqueous media due to the unalterable hydrophobicity.29 Novel probes with unique lumogen mechanism suitable for the specific detection of H₂S in high aqueous contents are thus still highly desirable.

Aggregation-induced emission (AIE), first reported in 2001, is a unique phenomenon that propeller-shaped fluorophores are non-emissive when molecularly dissolved but are highly fluorescent upon aggregation.29 It is proposed that these molecules cannot pack through a π π stacking process due to the propeller shape in the aggregate state, which restricts the intramolecular rotation, blocks the non-radiative pathway and opens up the radiative channel, leading to aggregation-induced emission.28 AIE is particularly attractive for biological aqueous environment and have showed diverse applications in cell imaging and bioprobes.30-33 However, AIE-based H₂S-targeting probes are still unavailable. Here in this work, we disclose our related research efforts which resulted in the first probe with AIE characteristics for the sensing of biological H₂S. The probe, namely AIE-S, demonstrated high specificity and sensitivity. Its feasibility to monitor H₂S in live cells was also confirmed.

**Experimental section**

**Materials**

Dry ethanol was distilled from sodium prior to use. Water was purified by a Millipore filtration system. Other chemicals and reagents were from commercial supplies and used without further purification. Reactions were monitored by thin-layer chromatography (TLC) carried out on Silica gel 60 F254 plates supplied by Qingdao Puke Separation Material Corporation using UV light as the visualizing agent. Flash column chromatography was performed using 200-300 mesh silica gel supplied by Qingdao Marine Chemical Factory, Qingdao, China.

**Instruments**

1H and 13C NMR spectra were recorded on a Bruker Fourier transform 400 NMR spectrometer at 25°C using CDCl 3 as solvent and were calibrated referencing residual undeuterated solvent as an internal reference (1H NMR = 7.26, 13C NMR = 77.16). All chemical shifts were given in ppm and coupling constants (J) in Hz. The following abbreviations or combinations thereof were used to explain the multiplicities: s = singlet, t = triplet, m = multiplet. IR spectra were recorded on a Bruker Vector 22 spectrophotometer as KBr pellets. High resolution mass spectra (HRMS) were recorded on an Agilent 6224 TOF LC/MS spectrometer using ESI-TOF (electrospray ionization-time of flight). Absorption spectra were acquired using a Hitachi U-3010 spectrophotometer. Fluorescence measurements were carried out on a Perkin-Elmer LS 45 fluorescence spectrometer. Dynamic light scattering (Beckman-Coulter Delsa™ Nano-C particle analyzer) was used to determine the size of the aggregated particle colloids in aqueous solution.

**Synthesis of AIE-S**

To a stirred solution of tetrphenylethene aniline (200 mg, 0.576 mmol) in dry ethanol was added salicylaldehyde (3.00 mL, 28.7 mmol). The reaction was allowed going at ambient temperature under N₂ atmosphere for 1 h. A yellow precipitate was formed gradually which was isolated by filtration and further purified by flash column chromatography on silica gel (petroleum ether/CH₂Cl₂, 5:1) to afford AIE-S as a yellow powder (234 mg, 90% yield). Ref = 0.45 (10:1 petroleum ether: EtOAc). 1H NMR (400 MHz, CDCl₃), δ (ppm): 13.30 (1 H, s), 8.59 (1 H, s), 7.37 (2 H, t, J 7.0), 7.15-7.01 (20 H, m), 6.93 (1 H, t, J 7.4). 13C NMR (100 MHz, CDCl₃), δ (ppm): 162.07, 161.29, 146.43, 143.75, 137.32, 143.67, 142.87, 141.54, 140.29, 133.16, 132.57, 132.30, 131.50, 131.47, 131.45, 127.98, 127.90, 127.80, 126.76, 126.73, 126.66, 120.68, 119.38, 119.15, 117.36. IR
Fluorometric analysis

N-(2-hydroxyethyl)piperazine-N’-2-ethanesulfonic acid buffer (HEPES, 10 mmol, pH 7.4) was prepared with purified water and purged with nitrogen for 5 minutes before use. AIE-S was dissolved in THF to make a 200 μM stock solution, which was diluted to 20 μM with HEPES for measurements. CuCl₂ and Na₂S were dissolved in the deoxygenated HEPES to make a stock solution of 1.0 mM. The spectrometer slit width of excitation and emission was adjusted to 5 and 10 nm, respectively, for measurements. All fluorometric experiments were performed at room temperature in triplicate.

Cell imaging

HeLa cells were cultured in DMEM supplemented with 10% fetal bovine serum in a humidified atmosphere of 5% CO₂ at 37°C. For imaging experiments, exponentially growing cells (at a density of 20000-40000 cells per well, respectively) were seeded in 24-well plate. Cells were cultured at 37°C in a 5% CO₂ atmosphere for 24 h before they were exposed to reagents. After the staining steps as described in figure captions, the images were collected upon excitation using the corresponding filters for DAPI (blue).

Results and discussion

AIE-S was designed by fusing a Cu²⁺ chelating salen group which functions as a water solubility-adjusting group, to tetraphenylethene, a typical AIE fluorogen.28 The philosophy of our design is as follow. Existing as aggregates, AIE-S in aqueous solution would be highly fluorescent, while treating AIE-S with Cu²⁺ would lead to fluorescent quenching due to the improved water solubility of the AIE-S-Cu²⁺ complex. However, ascribed to the high binding affinity between S²⁻ and Cu²⁺, Cu²⁺ would be de-coordinated from the complex in the presence of S²⁻ to release the less miscible AIE-S and therefore trigger the restoration of the luminescent property of the system due to the increased local concentration and thereof the aggregation of AIE-S (Fig. 1).

The synthesis of AIE-S was straightforward via the reaction between salicylaldehyde and tetraphenylethene aniline with the latter being prepared according to reported procedures.34 It was obtained as a yellow powder in a yield of 90% and characterized by NMR and high-resolution mass spectroscopy.

With AIE-S in hand, we first investigated its photoluminescent property. AIE-S is soluble and thus non-emissive in tetrahydrofuran (THF). Its solubility decreased gradually as the volume fraction of HEPES buffer (10 mM, pH = 7.4) increased. The system reached its threshold when the buffer fraction increased to 70%. In the THF-HEPES (3:7, v/v) mixture, AIE-S was still weakly fluorescent even though the system showed an average particle size diameter of 214.0 nm. When the buffer fraction was greater than 70%, the system exhibited obvious emission peaked at 533 nm which intensified as the buffer fraction increased (Supplementary Material: Fig. S1). These results indicated that AIE-S was a typical fluorogen with aggregation-induced emission characteristics.

Next, the fluorescent responses of AIE-S towards various concentrations of Cu²⁺ were studied. Experiments were carried out in HEPES buffer (10 mM, pH = 7.4, 25°C, 10% THF). As shown in Fig. 2, the blank solution of AIE-S (20 μM in the buffer) exhibited intense fluorescence peaked at 533 nm, while the addition of increasing concentrations of Cu²⁺ gradually attenuated the fluorescent signal and 95% quenching was observed when 0.5 equiv. Cu²⁺ was added. The coordination between Cu²⁺ and AIE-S was confirmed by UV-Vis absorption spectra analysis (Supplementary Material: Fig. S2). AIE-S showed an absorption band stretching from 320-470 nm with the maximum peak centered at 375 nm assigned to the lowest-lying π-π* transition. Treatment of AIE-S with Cu²⁺ shifted the absorption with two peaks at 330 and 400 nm, indicating the coordination of AIE-S with Cu²⁺. The stoichiometry of the complex was also determined to be 2:1 by a Job Plot indicating two AIE-S molecules bind to one Cu²⁺ ion (Fig. 3).
Fig. 2. The turn-off responses of AIE-S towards Cu²⁺. The addition of increasing concentrations of Cu²⁺ (0-20 μM) gradually quenched the fluorescence of AIE-S (20 μM). Spectra were taken in HEPES (10 mM, pH 7.4, 25°C, 10% THF) immediately after the addition of Cu²⁺ (λex 350 nm). (a) Spectra collected between 490-675 nm. (b) Correlation between the fluorescence intensity at 533 nm of the system (y) and the molar ratio of Cu²⁺ to AIE-S (x).

Fig. 3. Job’s plot for determining the stoichiometry of AIE-S and Cu²⁺ in HEPES (10 mM, pH 7.4, 25°C, 10% THF). The total concentration of AIE-S and Cu²⁺ was 20 μM and the excitation wavelength was 350 nm.

The selectivity of AIE-S to Cu²⁺ over other cations was then evaluated and it turned out that the fluorescent turn-off response of AIE-S was selective and specific for Cu²⁺. None of the other cations, including Ag⁺, Al³⁺, Ca²⁺, Cd²⁺, Co²⁺, Fe²⁺, Fe³⁺, Hg²⁺, K⁺, Mg²⁺, Mn²⁺, Na⁺, Ni²⁺, Pb²⁺, and Zn²⁺ was able to generate any obvious quenching effects. Moreover, the fluorescence of AIE-S could still be quenched by Cu²⁺ in the coexistence of competitive cations mentioned above, indicating that the coordination effect of AIE-S was specific towards Cu²⁺ (Fig. 4).

Fig. 4. Fluorescence responses of AIE-S towards various cations with or without the presence of Cu²⁺. Data were obtained as the fluorescence intensity at 533 nm (λex 350 nm) in HEPES buffer (10 mM, pH 7.4, 25°C, 10% THF) right after treating AIE-S with the cation indicated. The concentrations of AIE-S and the cations were both 20 μM. For the coexistence experiments, both Cu²⁺ and the tested substrates were kept at a final concentration of 20 μM.

Fig. 5. Size distribution of AIE-S (20 μM) before (a) and after (b) the addition of 0.5 eq Cu²⁺ in HEPES (10 mM, pH 7.4, 25°C, 10% THF). Data were determined by dynamic light scattering.

To confirm the mechanism shown in Fig. 1, dynamic light scattering measurements were carried out which showed that AIE-S probe has an average diameter of 249.0 nm in HEPES buffer (10% THF), while treating the probe with Cu²⁺ decreased the aggregation with the average diameter reduced to 169.2 nm (Fig. 5), supporting a plausible mechanism based on copper-initiated de-aggregation.

Having established AIE-S as a Cu²⁺ selective complexing ligand, we next study the ability of the AIE-S-Cu system to sense H₂S. For this purpose, Na₂S was employed as an aqueous sulfide source. Much to our delight, the addition of Na₂S to the weakly emissive AIE-S-Cu solution triggered an immediate increment in fluorescence intensity, demonstrating the release and aggregation of AIE-S (Fig. 6). The signal response towards Na₂S was instant and maximized within seconds since the fluorescent spectrum obtained right after the addition of Na₂S (t < 1 min) was overlapped with that of 30 minutes (Supplementary Material: Fig. S3). Moreover, the logarithm of the fluorescence intensity growth ratio of the system was linear to the concentration of Na₂S ranging from 2-15 μM with a correlation coefficient of 0.985 (Fig. 6c).

To examine if the fluorescence switch-on property of the AIE-S-Cu complex was specific towards sulfide, its responses towards various common anions and biothiols in HEPES buffer were measured. As shown in Fig. 7, in the presence of various anions, including bio-relevant F⁻, Cl⁻, Br⁻, I⁻, H₂PO₄⁻, NO₃⁻, CO₃²⁻, ClO₄⁻, and sulfur oxyanions such as SO₃²⁻, SO₄²⁻, S₂O₃²⁻, only S²⁻ was able to induce a significant fluorescence enhancement. Moreover, common biothiols such as glutathione (GSH) and L-cysteine (L-Cys) which are the major interferents jeopardizing the specificity of reported H₂S probes, proved to be innocent in this case (Fig. 7b). What is more, Na₂S could still induce a great fluorescence enhancement in the presence of the competitive anions. These results taken together demonstrated the high selectivity of the AIE-S-Cu complex for H₂S and its feasibility to detect H₂S in the complex biological samples.
Fig. 6. The fluorescence turn-on responses of AIE-S-Cu towards Na₂S. (a) Fluorescence spectra of AIE-S-Cu in the presence of various concentrations of Na₂S. Spectra were taken in HEPES buffer (10 mM, pH 7.4, 25°C, 10% THF) instantly after the addition of Na₂S (0-25 μM). The concentrations of AIE-S and Cu²⁺ were 20 and 10 μM, respectively (λex 350 nm). (b) Plot of fluorescence intensity at 533 nm vs concentrations of Na₂S. (c) Regression equation between the logarithm of the fluorescence intensity (mean ± SD, n = 3) growth ratio at 533 nm (y) and the corresponding concentration of Na₂S (x). I₀ was the fluorescence intensity of the quenched AIE-S-Cu.

Fig. 7. Fluorescence emission spectra of the AIE-S-Cu (20 μM for AIE-S and 10 μM for Cu²⁺) in HEPES buffer (10 mM, pH 7.4, 25°C, 10% THF) right after the addition of Na₂S or other biologically relevant anions, biothiols (20 μM) (λex = 533 nm). (a) Spectra collected between 490-675 nm. (b) Emission intensity at 533 nm (y) when different species (x) were added.

As a proof of concept we tested the probe in cellular imaging for the detection of sulfide. HeLa cells were incubated with the combination of AIE-S (5 μM) and CuCl₂ (3 μM) for 2 h. The cells were then rapidly washed with PBS and only weak fluorescence was observed under microscopy (Fig. 8a). The cells were then treated with Na₂S (100 μM) for 15 min at 37°C and resulted in a dramatic increase of intracellular fluorescence (Fig. 8b), which demonstrated the feasibility of AIE-S to image sulfide in live cells.

Conclusions

We have devised a novel fluorescent probe, AIE-S for the sensitive and rapid detection of H₂S in living cells. AIE-S featured with high selectivity, rapid response and aggregation-induced emission characteristics. The preparation of AIE-S is operationally simple and the light-up mechanism is based on S²⁻ initiated de-coordination and thereafter aggregation of the AIE-S-Cu complex. This probe enables monitoring of H₂S in solution and in cells with a high signal-to-noise ratio.

Supplementary Material

Scheme S1, Fig.S1 – S3, etc. http://www.thno.org/v04p1233s1.pdf

Competing Interests

The authors have declared that no competing interest exists.

References


